

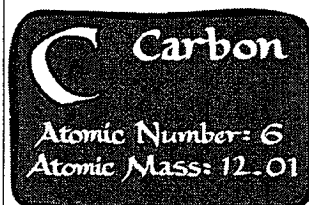
*Electrons = same as atomic number*

*Protons = same as atomic number*

*Neutrons =  
atomic mass - protons*

## Bohr Diagrams

- 1) Find your element on the periodic table.
- 2) Determine the number of electrons – it is the same as the atomic number.
- 3) This is how many electrons you will draw.



## Bohr Diagrams

PERIODS

- Find out which period (row) your element is in.
- Elements in the 1<sup>st</sup> period have one energy level.
- Elements in the 2<sup>nd</sup> period have two energy levels, and so on.

1st energy level can hold 2 electrons

2nd energy level can hold 8 electrons

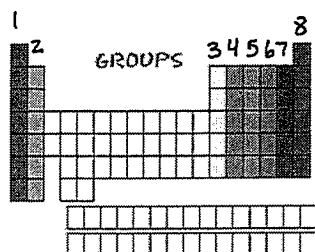
3rd energy level can hold 18 electrons

4th energy level can hold 32 electrons

5th energy level can hold 50 electrons

6th energy level can hold 72 electrons

## Lewis Structures



- Find out which group (column) your element is in.
- This will tell you the number of valence electrons your element has.
- You will only draw the valence electrons.

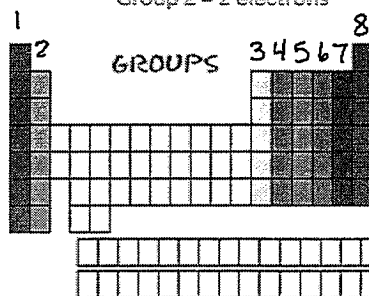
## Groups - Review

Group 8 = 8 electrons

Group 1 = 1 electron

Group 2 = 2 electrons

Except for He, it has 2 electrons



Each column is called a "group"

Each element in a group has the same number of electrons in their outer orbital, also known as "shells".

The electrons in the outer shell are called "valence electrons"

1. 6
2. Carbon
3. 12.001
4. P= 6 N=6 E=6



C